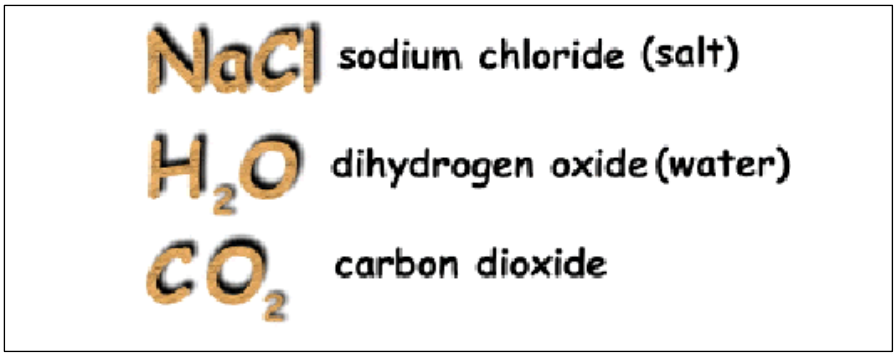
1. Compound

1. A pure _____ that contains _____ or more _____.

2. Chemical _____

2. Contains atomic _____ and subscripts to show the _____ and the _____ of atoms of each _____ in the compound.



3. What is a molecule?

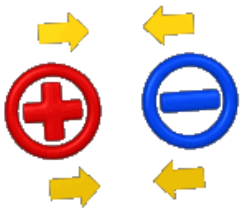
3. A _____ particle that forms as a result of _____ _____.

4. What is a CHEMICAL BOND?

4. A _____ that holds _____ _____ in a compound.

5. What is an _____ Bond?

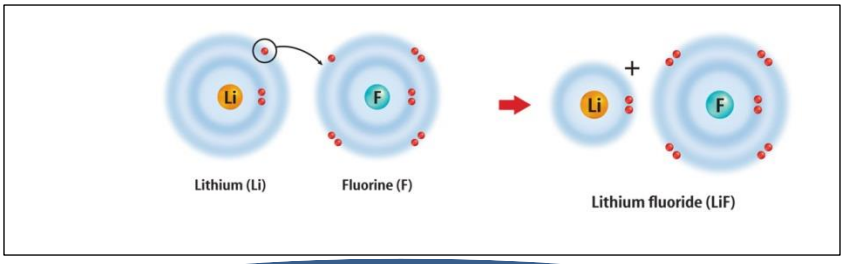
5. An _____ _____ between _____ and _____ charged _____ in an ionic compound.



- An atom can become _____ by _____ one or more _____ to another atom. Become an _____

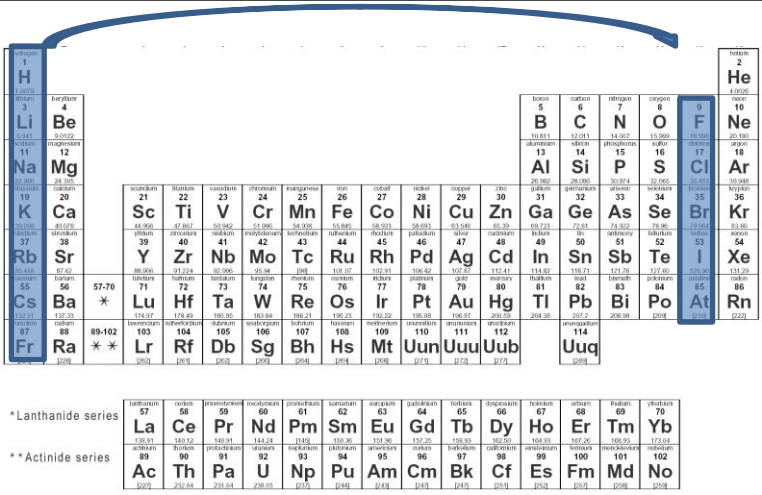
EXAMPLE:

- A lithium atom gives up an electron to a fluorine atom.
- The result is a positively charged lithium ion and a negatively charged fluoride ion.

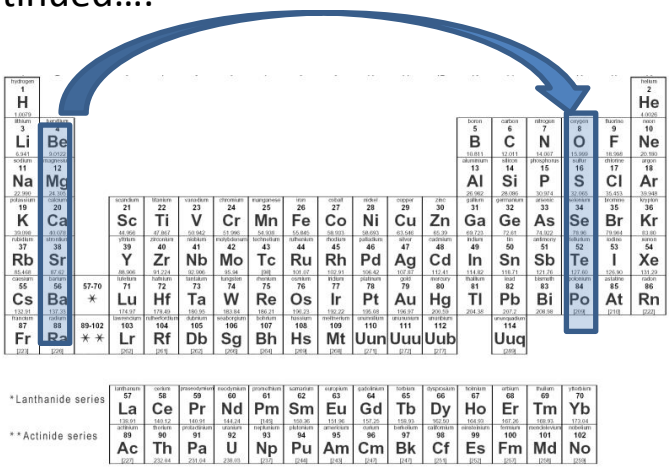


6. Likely to form bonds….

6.



6. continued….



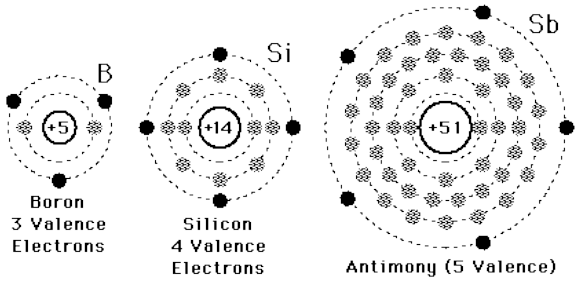
7. What are properties of _____ _____?

7.

- _____ and break apart easily
- Have _____ melting and boiling points
- Many _____ in water
- Usually _____ at room temperature

8. What are <u>VALENCE ELECTRONS?</u>

8. Electrons in the _____ _____ level.



9. What does a Lewis Dot Diagram look like?	9.
10.What kind of elements don't react?	10. _____ that are _____ rarely react to form compounds.
11.What is a _____ _____	11. A _____ bond formed when atoms _____ electrons. (potluck) • Elements that are _____ together on the periodic table are _____ likely to share electrons in a _____ bond than to _____ electrons.
12. What are _____ Compounds?	12. *Organic compounds are covalent compounds containing* _____ *atoms and are important for* _____ *organisms.*
13. What are properties of Covalent bonds?	13. • Can be solids, liquids, or gases at room temperature • Usually have _____ melting and boiling points than ionic compounds • Do _____ usually separate in water • Most do _____ conduct electricity

14. Single Bond	14. A bond formed when sharing _____ electron.
15. Double and Triple Bonds	15. A bond formed when sharing _____ or _____ electrons.