Complete and study the following for the upcoming test!!!

## Define and give examples of the following terms:

1) Protons:_____

2) Neutrons :_____

3) Electrons: :_____

4) Atomic Number:_____

5) Atomic Mass:_____

6) Isotope:_____

7) Atom:_____

8) Nucleus:_____

## Describe what the following people found/discovered and how:

**9)** Bohr
**10)** Rutherford
**11)** Dalton
**12)** Thomson
**13)** Lavosier

**Answer the following short answer questions that will be on the test on a separate sheet of paper:**

a. Thomson's model of the atom pictured electrons embedded in a ball of positive charge. Analyze why Rutherford's gold-foil experiment led to a change in this model.
b. The isotopes of carbon (carbon-12, carbon-13, carbon-14) have different mass numbers. Each isotope has six protons. Calculate the number of neutrons in each isotope.
c. Define the law of conservation of mass. Use an example to help define the law.
d. Table salt is made from a sodium ion, $Na^+$ and a chloride ion $Cl^-$. Represent with symbols how the two ions form to create a compound and explain how this occurs.
e. Be able to draw an atom from the periodic table with it's electron orbitals and nucleus!