

**Data: Table 1: Chemical Equations**

| <b>Make the following Equations on your desk</b>                                 | <b>Reactants</b> | <b>Products</b> | <b>Reactants - Final</b> | <b>Products - Final</b> | <b>Balanced Equation</b> |
|--|------------------|-----------------|--------------------------|-------------------------|--------------------------|
| $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$                         |                  |                 |                          |                         |                          |
| $\text{H}_2\text{O}_2 \rightarrow \text{H}_2\text{O} + \text{O}_2$               |                  |                 |                          |                         |                          |
| $\text{Na} + \text{O}_2 \rightarrow \text{Na}_2\text{O}$                         |                  |                 |                          |                         |                          |
| $\text{N}_2 + \text{H}_2 \rightarrow \text{NH}_3$                                |                  |                 |                          |                         |                          |
| $\text{P}_4 + \text{O}_2 \rightarrow \text{P}_4\text{O}_{10}$                    |                  |                 |                          |                         |                          |
| $\text{Fe} + \text{H}_2\text{O} \rightarrow \text{Fe}_3\text{O}_4 + \text{H}_2$  |                  |                 |                          |                         |                          |
| $\text{C} + \text{H}_2 \rightarrow \text{CH}_4$                                  |                  |                 |                          |                         |                          |
| $\text{Na}_2\text{SO}_4 + \text{CaCl}_2 \rightarrow \text{CaSO}_4 + \text{NaCl}$ |                  |                 |                          |                         |                          |
| $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ |                  |                 |                          |                         |                          |
| $\text{Al}_2\text{O}_3 \rightarrow \text{Al} + \text{O}_2$                       |                  |                 |                          |                         |                          |



